

Exp. 8 Report
Solubility of Ionic Precipitates

Name: _____

A. Solubility. Common Ion Effect

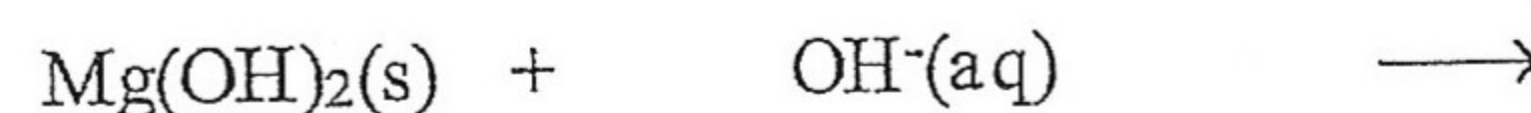
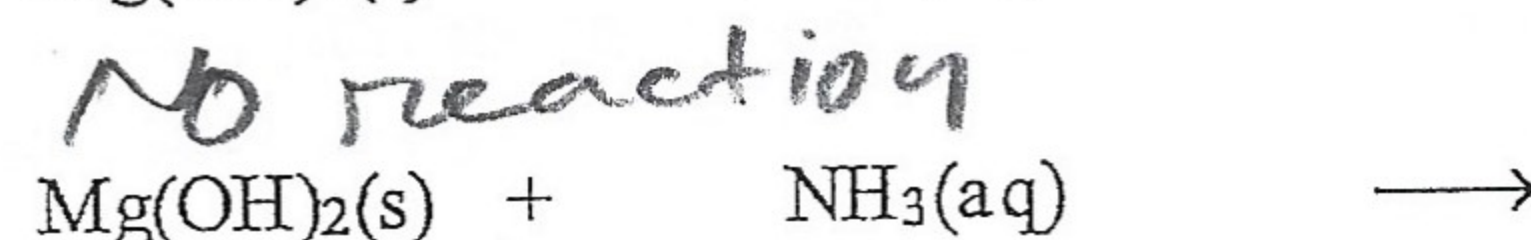
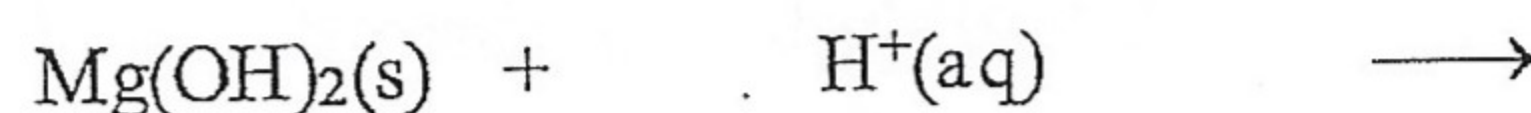
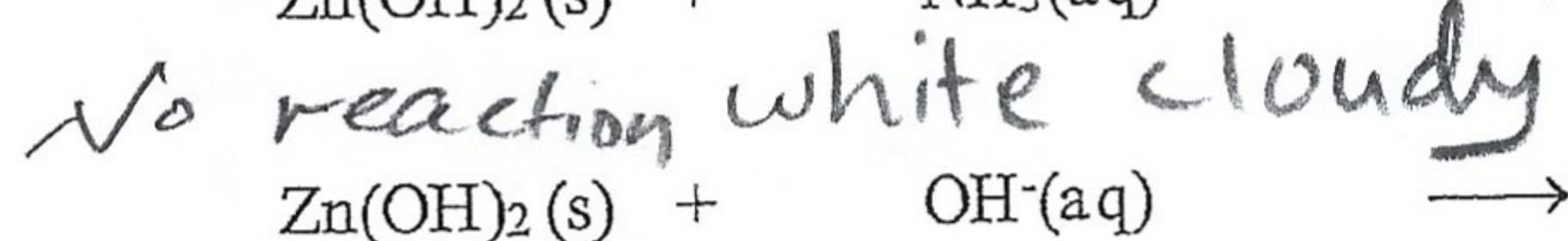
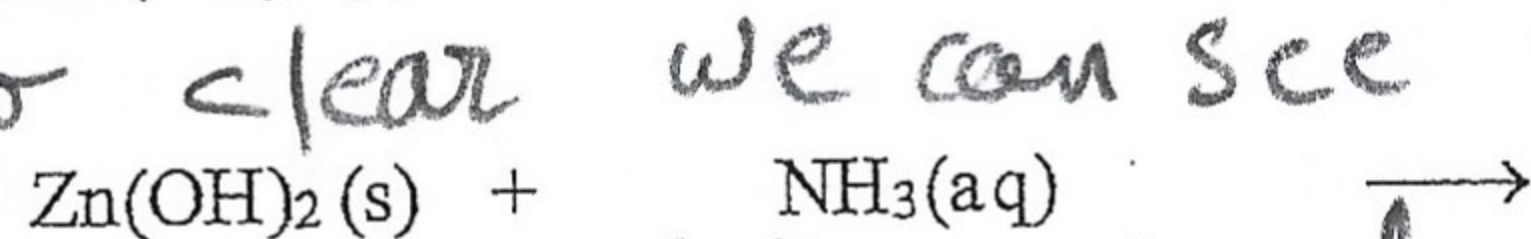
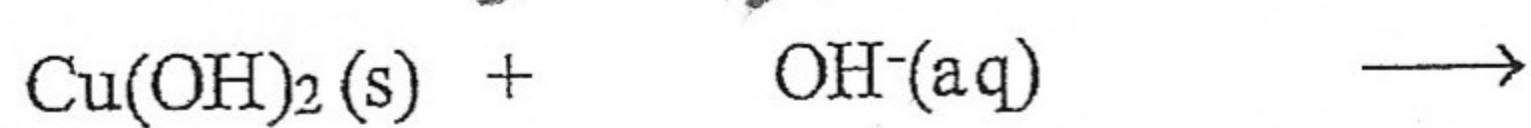
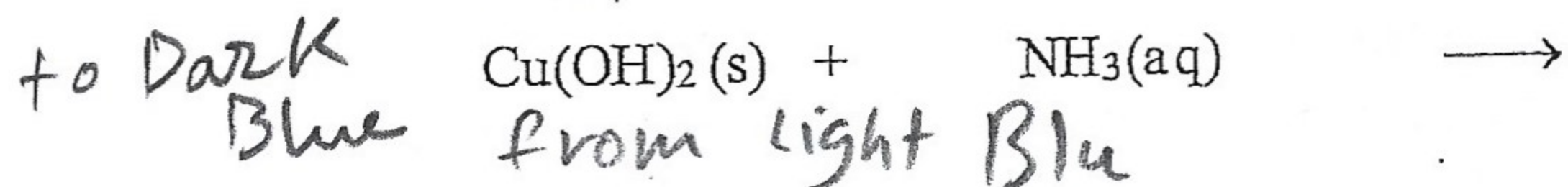
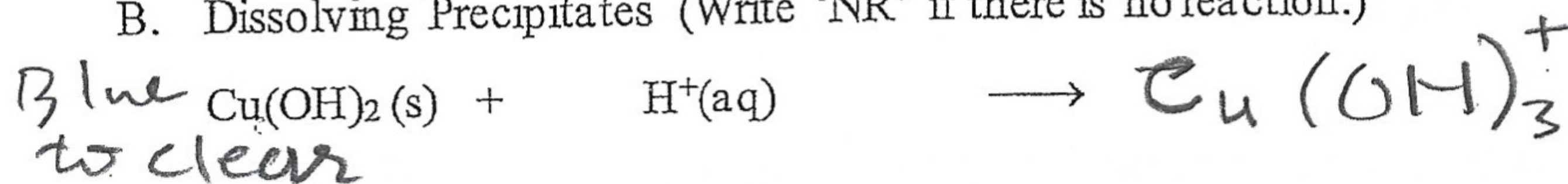
0.2 M Pb(NO ₃) ₂ drops	0.4 M HCl drops	V _f mL	[Pb ²⁺]	[Cl ⁻]	s mol/L	K _{sp}
5	5	1	0.035	0.07	0.035	1.72 × 10 ⁻⁵
10	5	2	0.07	0.07	0.035	0.000343
5	10	2	0.035	0.14	0.035	0.000686

0.035 (0.07)

_____ mL/drop for 0.200 M Pb(NO₃)₂

_____ mL/drop for 0.400 M HCl

B. Dissolving Precipitates (Write 'NR' if there is no reaction.)



Question

- From Part A, does solubility increase or decrease in the presence of common ions compared to solubility in pure water? Why?
- Tell what you observe when ammonia is added drop by drop to an aqueous solution of cupric sulfate.